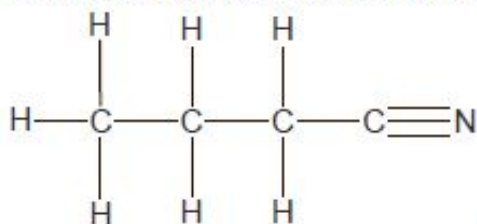


## Mark Scheme - 2.7 Alcohols and Carboxylic Acids

- 1 (a) (i) Nucleophilic substitution / Hydrolysis [1]
- (ii) Dissolved in alcohol (1) Propene or unambiguous structure (1) [2]
- (iii) Potassium manganate(VII) / Potassium dichromate(VI) - must be name (1)  
Oxidation (1) [2]
- (iv) (Add Potassium dichromate(VI)) and distil off the propanal from the reaction mixture [1]

- (b) (i) Step 1: Potassium cyanide in ethanol / Heat (1)  
Step 2: Heat with aqueous hydrochloric acid (or other acid) (1)



(1)

[3]

- (ii) Two points from different bullet points – 1 mark each.
- Atom economy / Amount of waste / Whether waste material was recyclable / Whether waste was toxic.
  - Amount of energy required / temperature required / pressure required / conditions used
  - Rate of production / time
  - Availability of catalyst
  - Cost of reactants / Availability of reactants / toxicity of reactants.
  - Two step processes usually have lower yields than one step processes / percentage yield [2]
  - Purification method / separation

- (c) (i) Butanoic acid is  $\text{C}_4\text{H}_8\text{O}_2$  so  $M_r = 88$  (1)  
Percentage carbon =  $48/88 \times 100 = 54.5\%$ ; percentage hydrogen =  $8/88 = 9.1\%$ ;  
Percentage oxygen =  $32/88 = 36.4\%$  (At least two of these for 1)  
OR empirical formula for butanoic acid =  $\text{C}_2\text{H}_4\text{O}$  (1) and  
calculate empirical formula from percentage masses =  $\text{C}_2\text{H}_4\text{O}$  (1) [2]

(ii) Structure 1 mark + 4 marks for explanations.

- Product is ethyl ethanoate. (1)
- Two points from the following required for each mark– MAX 4 marks
  - Sweet-smelling = ester
  - Peak at 1.0ppm implies – CH<sub>3</sub>
  - Peak area 3 = CH<sub>3</sub>
  - Peak area 2 = CH<sub>2</sub>
  - Triplet shows CH<sub>3</sub> is next to a CH<sub>2</sub> group.
  - Singlet shows CH<sub>3</sub> no hydrogen atoms bonded to adjacent carbon.
  - Peak at 2.1 ppm suggests this is next to C=O.
  - Quartet shows CH<sub>2</sub> is adjacent to a CH<sub>3</sub> group.
  - Peak at 4.0 ppm shows it is –O-CH<sub>2</sub>-
  - IR Peak at 1752 cm<sup>-1</sup> = C=O
  - IR Peak at 2981 cm<sup>-1</sup> = C-H or O-H
  - Cannot be –OH as we know there is no –OH in NMR spectrum

[5]

*QWC: selection of a form and style of writing appropriate to purpose and to complexity of subject matter. (1)*

*QWC: organisation of information clearly and coherently; use of specialist vocabulary where appropriate. (1)*

[2]

[20 marks]

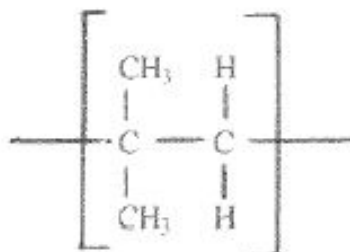
2.

(a) (i) (2-)Methylpropan-2-ol [1]

(ii) 30.1 / 30 [1]

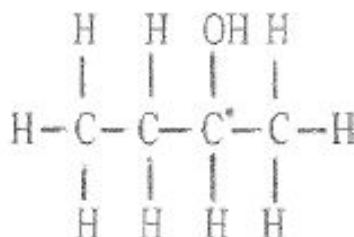
(iii) (Concentrated) sulfuric acid / phosphoric acid / aluminium oxide / pumice [1]

(iv)



(with or without n) [1]

(v)



(1) for structure, (1) for asterisk [2]

(vi) I acidified potassium dichromate /  $\text{H}^+$ ,  $\text{Cr}_2\text{O}_7^{2-}(\text{aq})$  [1]

II ethanal has a  $\text{C} = \text{O}$  bond at  $1650\text{-}1750\text{ cm}^{-1}$   
(metaldehyde does not have this bond) (1)

metaldehyde has a  $\text{C} - \text{O}$  bond at  $1000\text{-}1300\text{ cm}^{-1}$   
(ethanal does not have this bond) (1) [2]

(b) (i) Reagent 2,4-dinitrophenylhydrazine / 2,4-DNP OR iodine / NaOH or KI / NaOCl (1)

Observation yellow / orange / red precipitate OR yellow precipitate (1) [2]

(ii) Reagent ethanol / sulfuric acid OR  $\text{NaHCO}_3$  OR  $\text{Ag}^+/\text{NH}_3$  / Tollens' (1)

Observation sweet smelling liquid OR effervescence OR silver mirror (1) [2]

Total [13]

3. (a) (i)  $2\text{C}(\text{s}) + 3\text{H}_2(\text{g}) + \frac{1}{2}\text{O}_2(\text{g}) \rightarrow \text{C}_2\text{H}_5\text{OH}(\text{l})$  (state symbols needed) [1]  
 C(s) allowed as C(gr) or C(graphite) [1]
- (ii) (if these elements were reacted together) other products would form/ carbon does not react with hydrogen and oxygen under standard conditions [1]
- (b) (i) energy =  $100 \times 4.2 \times 54 = 22680$  [1]
- (ii) moles ethanol =  $0.81/46 = 0.0176$  (1)  
 energy change =  $\frac{22.68}{0.0176}$   $\Delta\text{H} = -1290$  (1)  
 -ve sign and correct to 3 sf (1) [3]
- (c) internet value numerically larger (1)  
 heat losses / incomplete combustion / thermal capacity of calorimeter ignored (1) no credit for energy loss [2]
- (d) (i)  $\text{C}_3\text{H}_7\text{OH} + 4\frac{1}{2}\text{O}_2 \rightarrow 3\text{CO}_2 + 4\text{H}_2\text{O}$  (ignore state symbols) [1]
- (ii) negative enthalpy change means energy in bonds broken is less than that in bonds made [1]
- (iii) more bonds broken and made in propanol and therefore more energy released [1]
- (e) any 4 from:  
 both conserve carbon / non-renewable fuel sources / fossil fuels / use renewable sources  
 (these gas / liquid) suitable for different uses e.g. ethanol to fuel cars  
 atom economy gasification is less (some C lost as  $\text{CO}_2$ ) /  $\text{CO}_2$  produced in gasification is a greenhouse gas  
 CO is toxic  
 gasification at high temperature / enzymes need low temperature  
 enzyme approach therefore saves fuel / gasification needs more energy [4]  
 3 max if any reference to destruction of ozone layer  
 QWC [2]  
 The candidate has selected a form and style of writing that is appropriate to purpose and complexity of the subject matter (1)  
 Answer has suitable structure (1)

**Total [17]**

4.

- (a) (i) ultraviolet / sunlight [1]  
(ii) A species with an unpaired electron. [1]
- (b)  $\text{CH}_4 + \text{Cl}\cdot \rightarrow \text{CH}_3\cdot + \text{HCl}$  (1)  
 $\text{CH}_3\cdot + \text{Cl}_2 \rightarrow \text{CH}_3\text{Cl} + \text{Cl}\cdot$  (1) [2]
- (c) (i) Two  $\text{CH}_3\cdot$  radicals combine (in a termination reaction). [1]  
(ii)  $24.3 \div 12 = 2.025$  for C  $4.1 \div 1.01 = 4.059$  H  $71.6 \div 35.5 = 2.017$  Cl (1)  
 $\text{CH}_2\text{Cl}$  (1) [2]
- (d) (i) Nucleophilic substitution [1]  
(ii) Methanol has hydrogen bonding between molecules (1)  
Chloromethane has van der Waals forces / dipole-dipole forces between molecules (1)  
Hydrogen bonding is stronger than Van der Waals/dipole-dipole (1) [3]  
(iii) Acidified potassium dichromate / acidified potassium manganate(VII) (1)  
Heat /warm (1) (Need correct reagent to gain heat mark) [2]
- (e) Compounds B and C are stable enough to reach the ozone layer OR Compound D would not reach the ozone layer as it would decompose in the lower atmosphere. (1)  
  
(The C-Cl forms)  $\text{Cl}\cdot$  which will decompose the ozone. (1)  
  
Compound A does not contain chlorine, (so it cannot form  $\text{Cl}\cdot$ ) / Compound A has a lower RODP (1) [3]

Total [16]

5.

- (a) (i)  $\delta^-$  on Br and  $\delta^+$  on C attached (1)  
Arrow from lone pair on  $\text{OH}^-$  to  $\delta^+$  on C (1)  
Arrow from C-Br bond to Br (1)  
Correct alcohol +  $\text{Br}^-$  (1) [4]
- (ii) Nucleophilic substitution [1]
- (iii) The bond breaks and both the electrons go to one of the bonded atoms/ the bond breaks and ions are formed. [1]
- (b) (i) Sodium hydroxide in ethanol/ alcohol [1]
- (ii) Elimination/ dehydrohalogenation [1]
- (iii) Structural formulae for but-1-ene (1)  
and but-2-ene (1) [2]
- (c) A is non-miscible with water/ does not mix with water and B is miscible/ mixes with water/ is soluble in water (1)  
A has a longer carbon chain/ is bigger (1)  
Hydrogen bonding (1)  
Between the OH in alcohol and water (1)  
In large alcohols non-polar/ hydrophobic part of molecule is large / OH is less significant part of molecule (1) [5]  
*QWC: organisation of information clearly and coherently; use of specialist vocabulary such as intermolecular force/ hydrogen bond/ hydrophobic/ non-polar/ miscible* [1]

**Total [16]**

6.

- (a) diagram completed with at least 1 water molecule and indication of interaction between O on one molecule and H on the other (1)

interaction between  $\delta^+$  on H and lone pair on O (1)

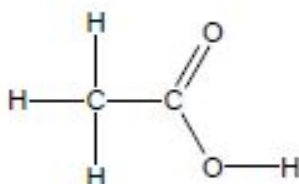
interaction labelled hydrogen bond (1) [3]

- (b) (i) reduction/ redox – accept 'oxidation' [1]

(ii) I OH [1]

II OH is also present in water [1]

- (c) (i) [1]



- (ii) peak at 1650-1750 (1)

due to C=O (1) [2]

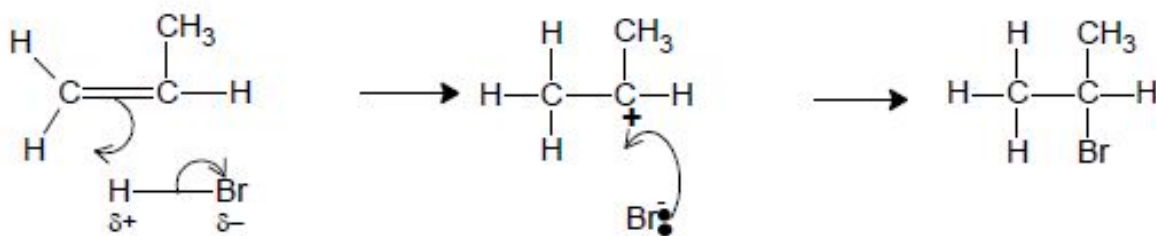
Total [9]

7. Reagent: acidified potassium dichromate / Cr<sub>2</sub>O<sub>7</sub><sup>2-</sup> and H<sup>+</sup>  
or acidified manganate(VII) / MnO<sub>4</sub><sup>-</sup> and H<sup>+</sup> (1)

Colour change: from orange to green  
or from purple to colourless (1) [2]

8.

- (a) (i) Molecules with different numbers of carbon atoms have different boiling points. [1]
- (ii) Any suitable reaction, e.g.  $C_{10}H_{22} \rightarrow C_4H_{10} + C_6H_{14}$  [1]
- (b) (i) Turns from orange to colourless (no credit for 'red') [1]
- (ii) (1) for arrows in first diagram; (1) for arrow in second diagram; (1) for all charges.



- (iii) Ethanol OR Alcohol solution / Heat - both required [1]
- (c) (i) Restricted rotation about double bond in but-2-ene but not butane (1)
- 2 groups attached to each carbon of the double bond are different in but-2-ene but in propene one carbon has the same two groups attached (1) [2]

(ii)



Accept any valid representation [1]

- (d) (i) Steam, phosphoric acid catalyst, (1) 300°C, 70 atm pressure (1) [2]
- (ii) Butan-2-ol will have IR absorptions at 2500-3550  $cm^{-1}$  / 1000 – 1300  $cm^{-1}$  and butene will not  
OR  
But-2-ene will have an IR absorption at 1620-1720 and butan-2-ol will not [1]

Total [13]



9.

- (a) Hydrogen bonding occurs between (1) oxygen, nitrogen or fluorine (1) of one molecule and hydrogen, which is bonded to oxygen / nitrogen / fluorine of another molecule (1)  
Alkanes do not contain an O-H, N-H or F-H bond and cannot therefore hydrogen bond to water molecules (1) [4]

QWC *Candidates should have use 'a selection and form of writing appropriate to purpose and to complexity of subject matter'* [1]

- (b) (i) The (purified) petroleum is separated by heating (1) due to the different boiling temperatures of different fractions (1)  
OR the mixture is vaporised (1) and then condensed according to boiling temperatures (1) (as at the oil refinery) [2]

- (ii)  $\text{CuCl}_2 \quad \text{Cu} +2 \quad \text{CuCl} \quad \text{Cu} +1$  (1)  
(reduction occurs when) the oxidation number becomes less positive (1) [2]

- (c) (i) Same molecular formula but a different structural formula / structure [1]

- (ii) Both of the carbon atoms of the double bond have different atoms / groups bonded to them (1)  
There is no free rotation about the double bond (1) [2]

- (iii)  $M_r$  of compound A is 146.3 / 146 (1)

$$\text{Cost per mole is } \frac{146.3 \times 48 \times 100}{100 \times 73} = \text{£}96.20 \quad (1)$$

(Accept £96.00 per mole if  $M_r$  of 146 has been used) [2]

Total [14]